

## Balancing Practice and Answer Key

1.  $\underline{\hspace{1cm}} \text{Sb} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{SbCl}_3$  \_\_\_\_\_
2.  $\underline{\hspace{1cm}} \text{AgNO}_3 + \underline{\hspace{1cm}} (\text{NH}_4)_2\text{CrO}_4 \rightarrow \underline{\hspace{1cm}} \text{Ag}_2\text{CrO}_4 + \underline{\hspace{1cm}} \text{NH}_4\text{NO}_3$  \_\_\_\_\_
3.  $\underline{\hspace{1cm}} \text{C}_3\text{H}_8 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$  \_\_\_\_\_
4.  $\underline{\hspace{1cm}} \text{NaClO}_3 \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{O}_2$  \_\_\_\_\_
5.  $\underline{\hspace{1cm}} \text{Fe} + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{FeCl}_2 + \underline{\hspace{1cm}} \text{H}_2$  \_\_\_\_\_
6.  $\underline{\hspace{1cm}} \text{Al} + \underline{\hspace{1cm}} \text{H}_2\text{SO}_4 \rightarrow \underline{\hspace{1cm}} \text{Al}_2(\text{SO}_4)_3 + \underline{\hspace{1cm}} \text{H}_2$  \_\_\_\_\_
7.  $\underline{\hspace{1cm}} \text{CH}_2 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$  \_\_\_\_\_
8.  $\underline{\hspace{1cm}} \text{Na}_2\text{CO}_3 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{H}_2\text{CO}_3$  \_\_\_\_\_
9.  $\underline{\hspace{1cm}} \text{C}_2\text{H}_6 + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$  \_\_\_\_\_
10.  $\underline{\hspace{1cm}} \text{Pb}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{H}_2\text{S} \rightarrow \underline{\hspace{1cm}} \text{PbS} + \underline{\hspace{1cm}} \text{HNO}_3$  \_\_\_\_\_
11.  $\underline{\hspace{1cm}} \text{Zn} + \underline{\hspace{1cm}} \text{AgNO}_3 \rightarrow \underline{\hspace{1cm}} \text{Zn}(\text{NO}_3)_2 + \underline{\hspace{1cm}} \text{Ag}$  \_\_\_\_\_
12.  $\underline{\hspace{1cm}} \text{KClO}_3 \rightarrow \underline{\hspace{1cm}} \text{KCl} + \underline{\hspace{1cm}} \text{O}_2$  \_\_\_\_\_
13.  $\underline{\hspace{1cm}} \text{C}_4\text{H}_{10} + \underline{\hspace{1cm}} \text{O}_2 \rightarrow \underline{\hspace{1cm}} \text{CO}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$  \_\_\_\_\_
14.  $\underline{\hspace{1cm}} \text{NaBr} + \underline{\hspace{1cm}} \text{Cl}_2 \rightarrow \underline{\hspace{1cm}} \text{NaCl} + \underline{\hspace{1cm}} \text{Br}_2$  \_\_\_\_\_
15.  $\underline{\hspace{1cm}} \text{SrBr}_2 + \underline{\hspace{1cm}} (\text{NH}_4)_2\text{CO}_3 \rightarrow \underline{\hspace{1cm}} \text{SrCO}_3 + \underline{\hspace{1cm}} \text{NH}_4\text{Br}$  \_\_\_\_\_
16.  $\underline{\hspace{1cm}} \text{Mg}(\text{OH})_2 + \underline{\hspace{1cm}} \text{HCl} \rightarrow \underline{\hspace{1cm}} \text{MgCl}_2 + \underline{\hspace{1cm}} \text{H}_2\text{O}$  \_\_\_\_\_
17.  $\underline{\hspace{1cm}} \text{Ba}(\text{OH})_2 \rightarrow \underline{\hspace{1cm}} \text{BaO} + \underline{\hspace{1cm}} \text{H}_2\text{O}$  \_\_\_\_\_

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1.	$\underline{2}$ Sb + $\underline{3}$ Cl <sub>2</sub> → $\underline{2}$ SbCl <sub>3</sub>	synthesis
2.	$\underline{2}$ AgNO <sub>3</sub> + $\underline{1}$ (NH <sub>4</sub> ) <sub>2</sub> CrO <sub>4</sub> → $\underline{1}$ Ag <sub>2</sub> CrO <sub>4</sub> + $\underline{2}$ NH <sub>4</sub> NO <sub>3</sub>	double replacement
3.	$\underline{1}$ C <sub>3</sub> H <sub>8</sub> + $\underline{4}$ O <sub>2</sub> → $\underline{3}$ CO <sub>2</sub> + $\underline{2}$ H <sub>2</sub> O	combustion
4.	$\underline{2}$ NaClO <sub>3</sub> → $\underline{2}$ NaCl + $\underline{3}$ O <sub>2</sub>	decomposition
5.	$\underline{1}$ Fe + $\underline{2}$ HCl → $\underline{1}$ FeCl <sub>2</sub> + $\underline{1}$ H <sub>2</sub>	single replacement
6.	$\underline{2}$ Al + $\underline{3}$ H <sub>2</sub> SO <sub>4</sub> → $\underline{1}$ Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub> + $\underline{3}$ H <sub>2</sub>	single replacement
7.	$\underline{2}$ CH <sub>4</sub> + $\underline{3}$ O <sub>2</sub> → $\underline{2}$ CO <sub>2</sub> + $\underline{2}$ H <sub>2</sub> O	combustion
8.	$\underline{1}$ Na <sub>2</sub> CO <sub>3</sub> + $\underline{2}$ HCl → $\underline{2}$ NaCl + $\underline{1}$ H <sub>2</sub> CO <sub>3</sub>	double replacement
9.	$\underline{2}$ C <sub>2</sub> H <sub>6</sub> + $\underline{7}$ O <sub>2</sub> → $\underline{4}$ CO <sub>2</sub> + $\underline{6}$ H <sub>2</sub> O	combustion
10.	$\underline{1}$ Pb(NO <sub>3</sub> ) <sub>2</sub> + $\underline{1}$ H <sub>2</sub> S → $\underline{1}$ PbS + $\underline{2}$ HNO <sub>3</sub>	double replacement
11.	$\underline{1}$ Zn + $\underline{2}$ AgNO <sub>3</sub> → $\underline{1}$ Zn(NO <sub>3</sub> ) <sub>2</sub> + $\underline{2}$ Ag	single replacement
12.	$\underline{2}$ KClO <sub>3</sub> → $\underline{2}$ KCl + $\underline{3}$ O <sub>2</sub>	decomposition
13.	$\underline{2}$ C <sub>4</sub> H <sub>10</sub> + $\underline{13}$ O <sub>2</sub> → $\underline{8}$ CO <sub>2</sub> + $\underline{10}$ H <sub>2</sub> O	combustion
14.	$\underline{2}$ NaBr + $\underline{1}$ Cl <sub>2</sub> → $\underline{2}$ NaCl + $\underline{1}$ Br <sub>2</sub>	single replacement
15.	$\underline{1}$ SrBr <sub>2</sub> + $\underline{1}$ (NH <sub>4</sub> ) <sub>2</sub> CO <sub>3</sub> → $\underline{1}$ SrCO <sub>3</sub> + $\underline{2}$ NH <sub>4</sub> Br	double replacement
16.	$\underline{1}$ Mg(OH) <sub>2</sub> + $\underline{2}$ HCl → $\underline{1}$ MgCl <sub>2</sub> + $\underline{2}$ H <sub>2</sub> O	double replacement
17.	$\underline{1}$ Ba(OH) <sub>2</sub> → $\underline{1}$ BaO + $\underline{1}$ H <sub>2</sub> O	decomposition